

## Chemical Reaction Types Quiz

Identify each chemical reaction as:

synthesis (S)

decomposition (D)

single replacement (SR)

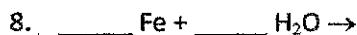
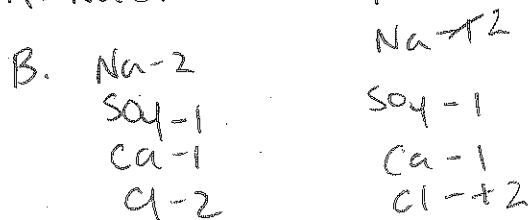
double-replacement (DR)

or combustion reaction (C)

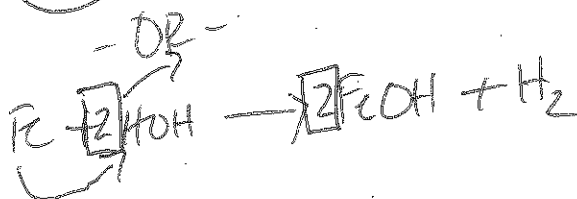
1. DR  $2 \text{Al}_2(\text{SO}_4)_3 + 3\text{Ca}(\text{OH})_2 \rightarrow 2\text{Al}(\text{OH})_3 + 3\text{CaSO}_4$
2. C  $2\text{C}_2\text{H}_2 + 5\text{O}_2 \rightarrow 4\text{CO}_2 + 2\text{H}_2\text{O}$
3. SR  $\text{Mg} + 2\text{AgNO}_3 \rightarrow \text{Mg}(\text{NO}_3)_2 + 2\text{Ag}$
4. DR  $3\text{Ba}(\text{NO}_3)_2 + 2\text{H}_3\text{PO}_4 \rightarrow \text{Ba}_3(\text{PO}_4)_2 + 6\text{HNO}_3$
5. D  $\text{Mg}(\text{ClO}_3)_2 \rightarrow \text{MgCl}_2 + 3\text{O}_2$
6. S  $2\text{Be} + \text{O}_2 \rightarrow 2\text{BeO}$

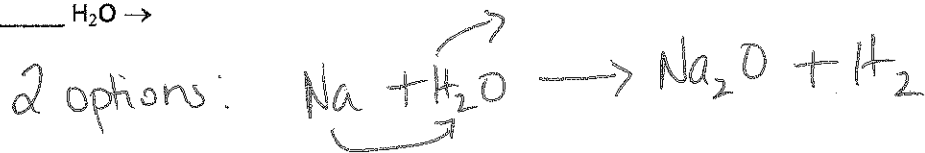
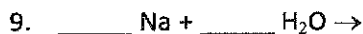
### For 7-11

- A. Predict the products (remember to write the correct chemical formula)
- B. Balance the Reaction (**SHOW YOUR INVENTORIES!!!**)
- C. Box in the coefficients.

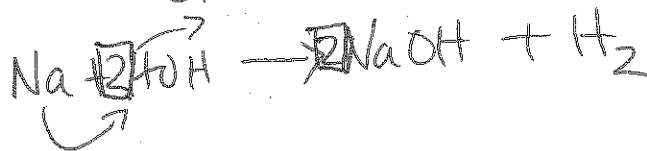


use  $\text{Fe}^{+1}$

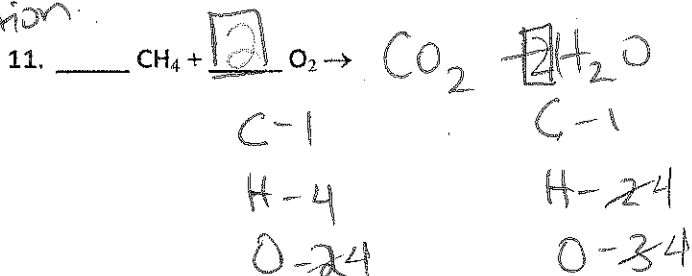




- OR -



Combustion:



12. Write the word equation for the following chemical reaction. (Don't forget about the states of matter):



Solid magnesium and aqueous iron(III) chloride react to form solid iron and aqueous magnesium chloride

13. When solid sodium reacts with aqueous magnesium iodide, the products are aqueous sodium iodide and solid magnesium. Write the chemical reaction using chemical formulas, note the states of matter in parentheses, and balance the equation.

